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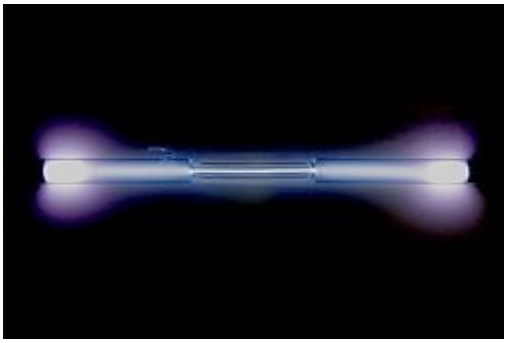
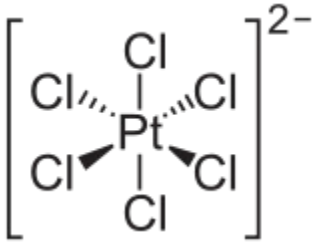
Reaction with the Noble Gasses



by Miles Mathis

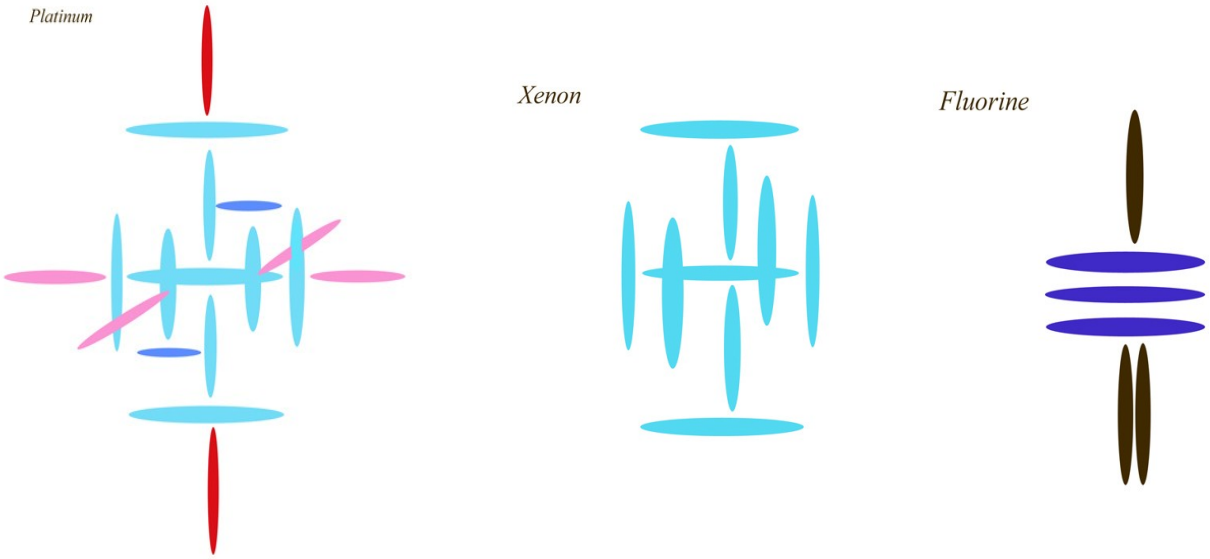
This paper addresses the third of the five "impossible chemisty" articles from *New Scientist*, January 23, 2012. After decades of believing the noble gasses were completely unreactive, Neil Bartlett discovered in 1961 that Xenon would react with Platinum hexaFluoride. That's fine, and I have nothing to say against it here. But let us use my nuclear diagrams to show *why* this is happening. I will show it has nothing to do with electrons. We could say it has to do with ionization potentials, but we will see that those potentials have to be assigned to the charge field being channeled through the nucleus, not to electrons orbiting it.

I have already diagrammed all the noble gasses, so we will begin by diagramming Platinum and PtF_6. Notice that my diagrams are a bit more advanced than the current ones. This is the kind of thing you get now:



Not so helpful.

In my diagrams, the first thing we notice is that Platinum, like Xenon, has a center made of triple alphas. That is what the blue-green disks represent. So we can plug a maximum of six protons in each hole. Platinum already has three protons in four holes (purple), and four protons in the other two (red).



Xenon has none. Why doesn't Xenon have any? Because elements are normally created under pressure in stars or galactic cores. It takes a lot of force to put those level four protons in the holes. Why? Because the outer disks of the third level (see Xenon) are perpendicular to the external charge field. Those perpendicular disks act as walls, and external charge is resisted. The charge channeling through the noble gasses is weak. See my previous nuclear papers for more on this.

So to make Xenon reactive takes a strong external charge field. The ambient charge field isn't focused enough to do that. Even the charge field of most atoms isn't focused enough to do that. Even the powerful Fluorine, which has a very sharp and focused charge field around it, can't do anything with the weak charge through Xenon; and this is because Fluorine only has those two protons recycling charge through its axial level. Xenon has six perpendicular protons in all directions blocking the ingress of Fluorine. Fluorine has enough charge focus, but not enough charge strength to get through. We need an even more focused and powerful charge field to force a charge channel through Xenon. Let us see how PtF_6 does that.

We put Fluorines in all six outer holes of Platinum. That is where you see the purple and red disks. Where we have purple disks we have three open holes, and where we have red disks we have two open holes. So all the six places can easily take a Fluorine.

My critic will say, "But I thought you just said putting protons in level 4 holes takes huge amounts of pressure, as in stars." Yes, putting the first proton in does, which is why big elements have to built up from Xenon in stars. But molecules can be built like this because after the first protons are in, the charge channels are already built. The other protons going in that hole will be much easier to put in because they are fitting into an existing charge channel. The stream is already running, as it were.

"So I should just be able to pop an extra proton into Platinum, making it into Gold? Gee, alchemy was never so easy!" Well, that is a good point, but it has an answer. These charge channels have varying strengths, and the charge channel of Hydrogen is very weak compared to the charge channel of the 4th level. The stream flowing through Hydrogen is moving much slower than the stream flowing through

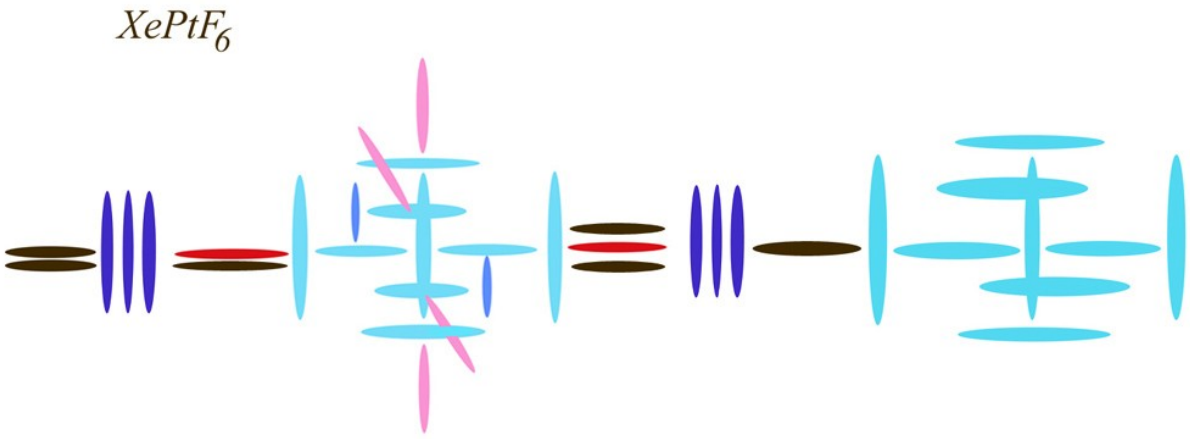
Platinum, and the streams don't fit together. Hydrogen bonds with smaller elements that have flow rates more in its range. So if we throw a single proton into that charge field around Platinum, it gets bounced out, like a canoe in a stream too big for it. I could create math and fancy names to explain the mechanics here, but for now I will stick with the visuals.

We also have to consider that Hydrogen and Helium have very unfocused charge channels. They are emitting charge in a full circle, basically, so the charge stream is not linear. Larger elements create channels that could be called more or less linear. They are more focused. This is one of the first things that stands in the way of easy alchemy. If Hydrogen and Helium had linear charge channels, alchemy would indeed be quite simple, as you put it.

"OK, so when Hydrogen bonds with something like Chlorine, why doesn't this just create Argon?" Because the structure of the nucleus is completely different, and the structure of the charge channels is completely different. Yes, they both have an atomic number of 18, but an element is more than an atomic number. An element is a certain nuclear structure.

"But Hydrogen can bond to Fluorine, and Fluorine can bond to Platinum, so the streams can't be that different." Another good point, but again there is an answer. One Fluorine *won't* bond to one Platinum, and that is one reason why (another reason is that it creates an unbalanced molecule). It takes six Fluorines to match the charge stream of Platinum, and one just gets bounced out. Yes, I am simplifying these answers because I want to move on, but I want to suggest that all such questions have mechanical answers. We just have to look for the answers in the charge channels, not in the electron orbitals.

Now, let us return to the question at hand. How does PtF_6 overcome the weak charge channels of Xenon? It does so by vastly increasing the charge channels of Fluorine alone. Once we have the Fluorines in all six slots, we have a huge amount of charge running through the PtF_6 molecule. I encourage you once again to think of the protons or alphas as fans pulling charge through the molecule. The more fans you have, the more charge you have moving through. Your charge channels will have a higher charge density, which just means that more charge photons are being swept through the nuclei. To make this even greater, we want a charge differential across the molecule, so we plug one Fluorine in the top with two prongs sticking out, and one Fluorine in the bottom with one prong sticking out. This makes it very clear to the charge field which direction it must go through the axial level. It goes north to south. We then plug the south end into the north end of the Xenon.



I have not drawn the carousel level Flourides of Platinum. With this diagram, we can see how PtF_6 is

able to force an entry into the top hole of Xenon. The *strength* of the charge channel is given us by the Platinum atom, which matches Xenon in its core. It is channeling charge with triple alphas, and this is what allows it to overcome the triple-alpha wall of resistance in that cap level disk of Xenon. But it also needs the *focusing* of the Fluorine atom on that end, which acts to sharpen the charge channel. As you see, we have the charge channel being funneled down from six protons as the charge leaves Platinum, to one proton as the charge leaves the Fluorine and goes into the Xenon. This is important, because the charge has to go into the hole. That top blue-green disk of Xenon is perpendicular to the charge channel, so it tends to block it. The only place that isn't blocking is the hole. That is the only parallel charge channel. So it is not beside the point that Fluorine is funneling the charge. Yes, those blue-green disks *can* take six protons, but that doesn't mean they always *want* to. The disks of the noble gasses *don't* want to, because they don't need to. They are already charge balanced, and so the channels have near zero potential. Xenon isn't *pulling* much charge to that hole, so if anything wants to get in, it has to do all the pushing itself. And it has to push at just the right place. Fluorine backed up by Platinum is able to do both those things simultaneously. It has both the charge strength and the charge focus to enter that hole.

Again, all this can be put into scientific language and equations later. The important thing is to start with some rational mechanics, and then build the math and fancy language around that. That is the opposite of what science has become, but in my opinion it is what science should be. Physics and chemistry have had a century to explain this, and they have done all they could with technical naming and pretty equations. It is way past time to give simple mechanics a try.